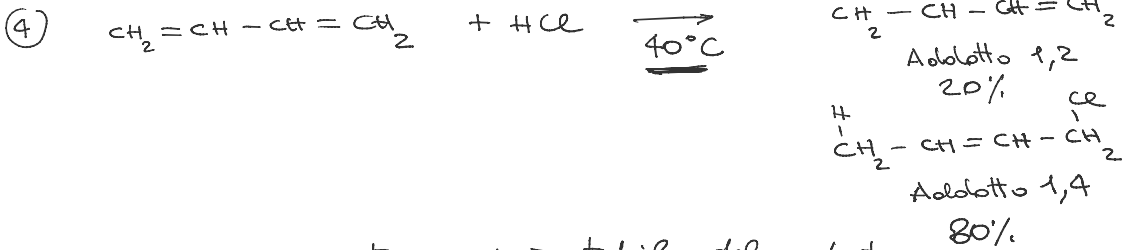
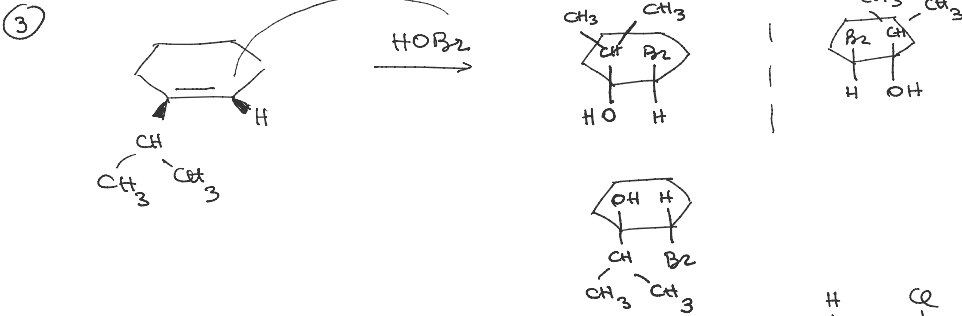
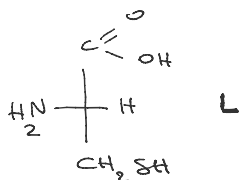
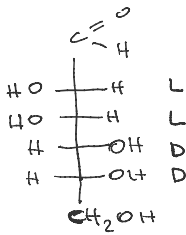
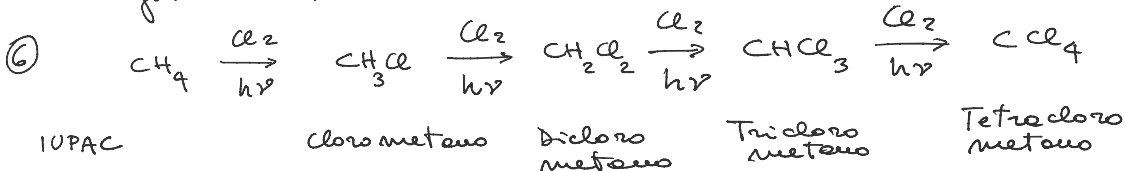
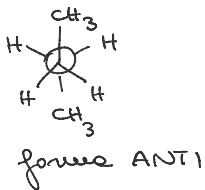


L or D?

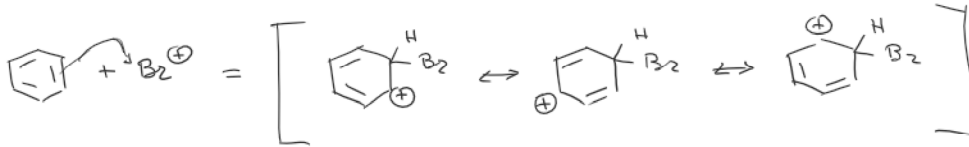


⑤ Newman per rotamero più stabile del n-butano

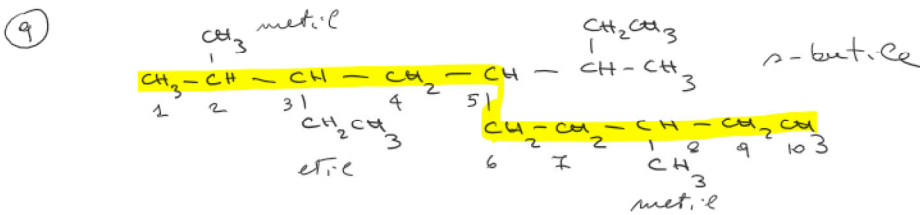
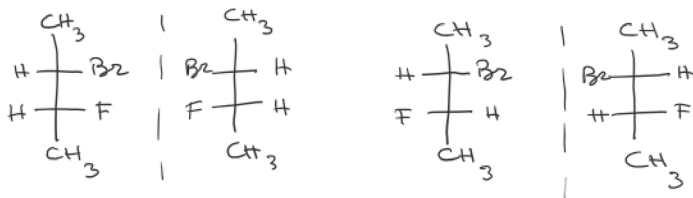
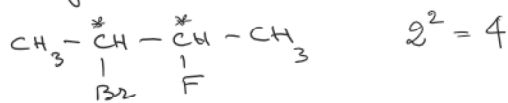


	1 Cl	2 Cl	3 Cl	4 Cl
IUPAC	Clorometano	Dicloro metano	Tricloro metano	Tetracloro metano
Comune	Cloruro di metile	Cloruro di metilene	Cloroformio	Tetracloruro di carbonio

7) Forme limite di risonanza per intermediario di Wheland nelle bromurazione.



8) Tutti gli stereoisomeri configurazionali del 2-bromo-3-fluorobutano utilizzando le formule di Fischer

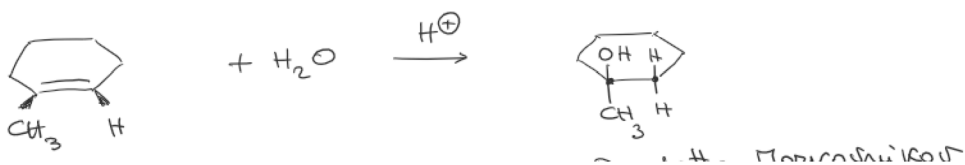


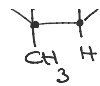
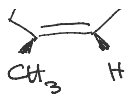
2,8-dimetil-3-etil-5-n-butil decano
 5-n-butil-3-etil-2,8-dimetil decano

Domande aperte

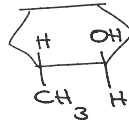
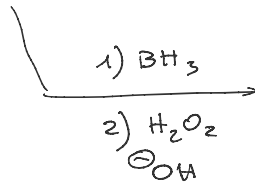
1) Polimerizzazione del propilene con meccanismo radicalico
 Polim. cationica e anionica
 Polimeri stereoregolari

2) Si descrivono i meccanismi di due reazioni che permettono di addizionare H₂O al 1-metil-1-cicloesene con orientamento opposto.





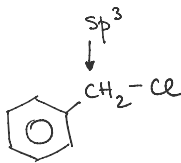
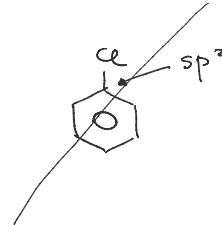
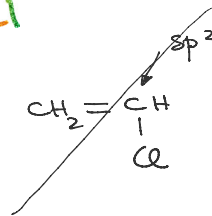
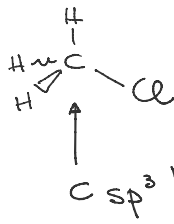
Prodotto Markovnikov



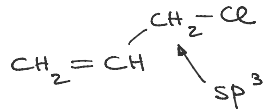
Prodotto Anti-Markovnikov

ALOGENURI ALCHILICI

R-X



Cloruro di benzile

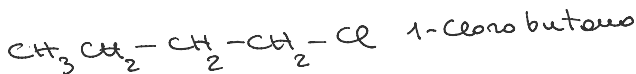


Cloruro di allile



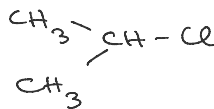
IUPAC
Bromoetano

COMUNE
Bromuro di etile



1-Clorobutano

Cloruro di butile



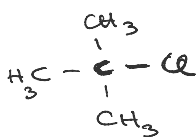
2-Cloropropano

Cloruro di isopropile

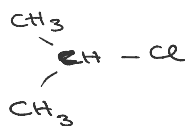
Utilizzi

- refrigeranti
- propellenti
- anestetici
- solventi (alcuni composti organici ma non sono infiammabili)
- importanti intermedi di sintesi

Classificazione



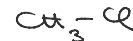
Alogenuro alchilico
TERZIARIO



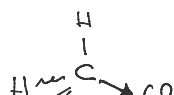
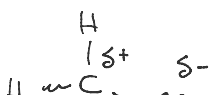
SECONDARIO

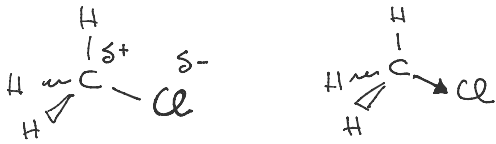


PRIMARIO



METILICO



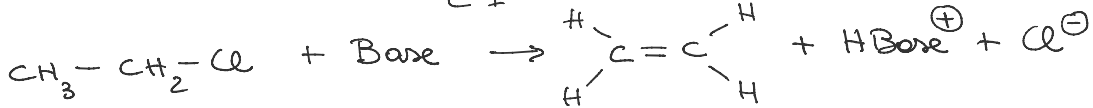


SOSTITUZIONE NUCLEOFILA ALCHILICA

S_N2 SOST. NUCL. BIMOLECOLARE MONOSTADIO
 S_N1 SOST. NUCL. MONOMOLECOLARE BISTADIO



ELIMINAZIONE $E2$ ELIMINAZIONE BIMOLECOLARE MONOSTADIO
 $E1$ ELIMINAZIONE MONOMOLECOLARE BISTADIO

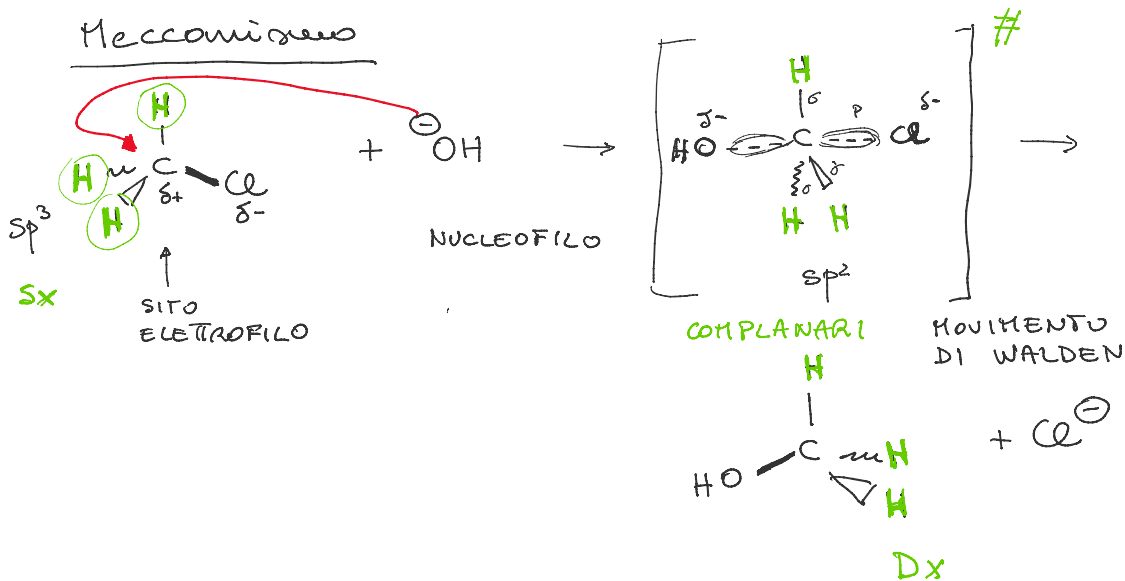


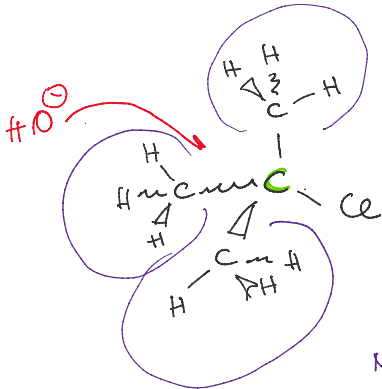
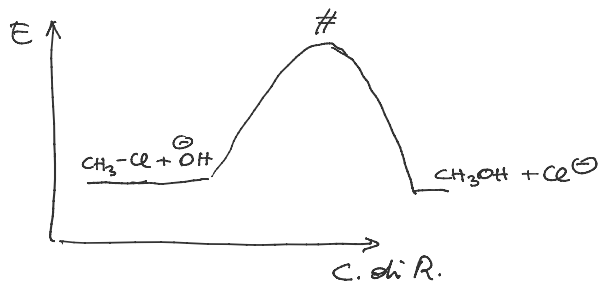
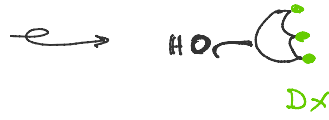
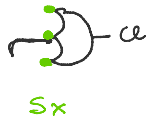
S_N2 Biomolecolare Monostadio

Evidenze sperimentali

- la reazione predilige substrati metilici/primari non ingombri stericamente
- Cinetica del 2° ordine complessivo
 $v = k [R-X]^1 [Nu^\ominus]^1$ $1+1=2$ ← ordine complessivo
 ↓
 velocità della reazione
- Stereochimica: la reazione avviene con INVERSIONE DI CONFIGURAZIONE
- Avviene a T base. ha un momento dipolo Non de-Ref a H
- Il solvente deve essere polare aprotico

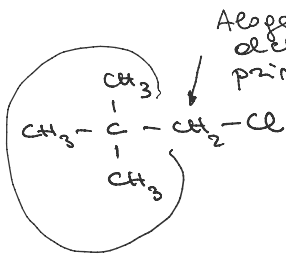
Meccanismi



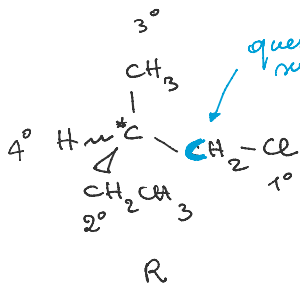
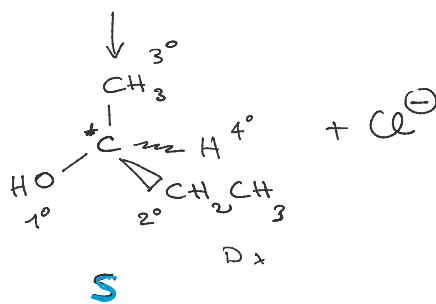
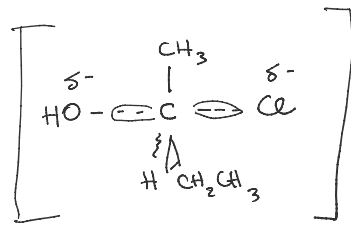
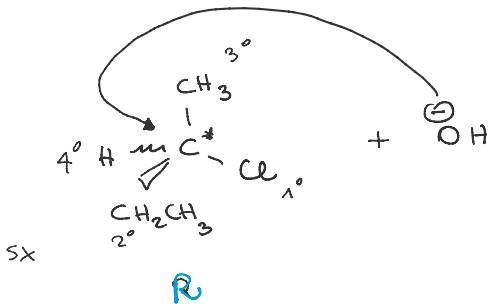


Notevole impedimento sterico all'entrata del Nu^\ominus da "dietro"

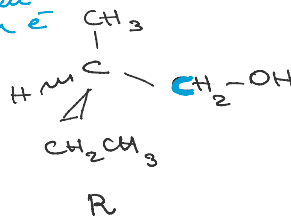
I substrati terziari non danno S_N2



INGOMBARO STERICAMENTE
 reagisce 100.000 volte più lentamente rispetto a CH_3Cl nelle S_N2

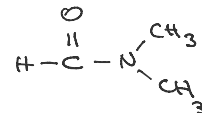


questo C subisce movimento di Walden ma non è chirale
 dunque non ne accorgo

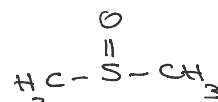


Solventi:

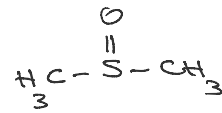
DMF Dimetilformamide



DMSO Dimetilossolformido



DMSO Dimetil solforio



(Solvente polare protico: H_2O Forme
leg. a parte H e "schermata" il
Nucleofilo
 $\ominus\text{OH}$)