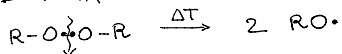


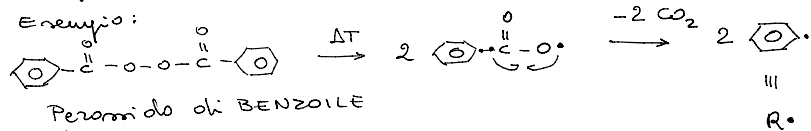
POLIMERI DI ADDIZIONE o A CRESCITA DI CATENA

MECCANISMO RADICALICO

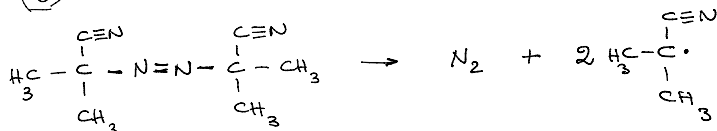
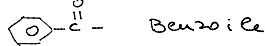
1. INIZIAZIONE



Esempio:



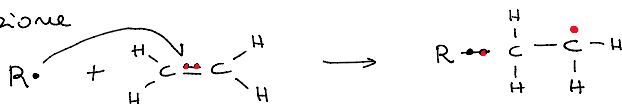
Perossido di BENZOILE



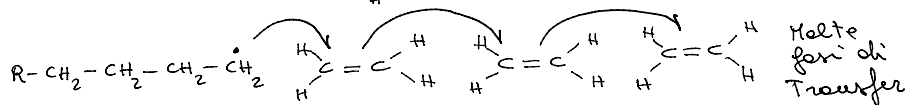
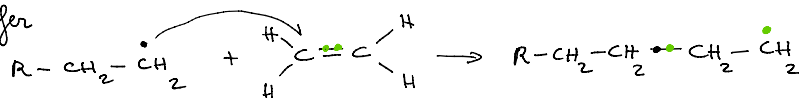
2,2'-Azobis(isobutyronitrile)

PROPAGAZIONE

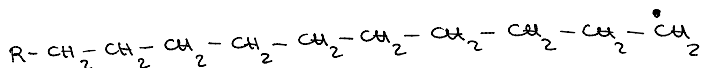
Addizione



Transfer

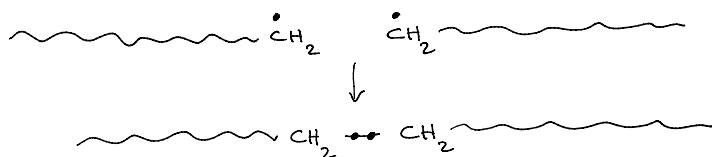


Molte
fasi di
Transfer

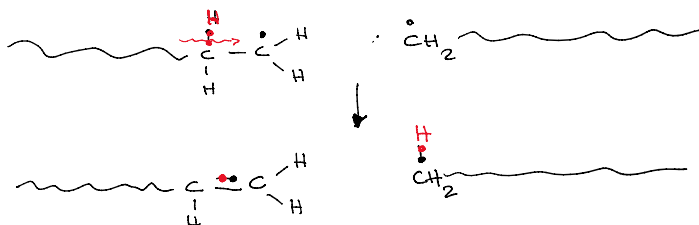


TERMINAZIONE

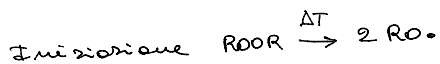
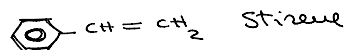
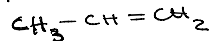
1. COMBINAZIONE



2. DISPROPORZIONE



Polymerizzazione radicalica del propilene



Propagazione

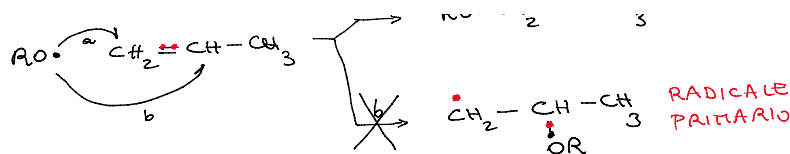
Addizione



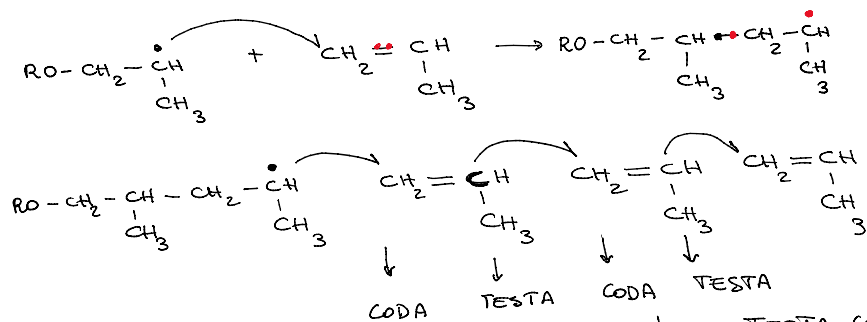
RADICALE
SECONDARIO



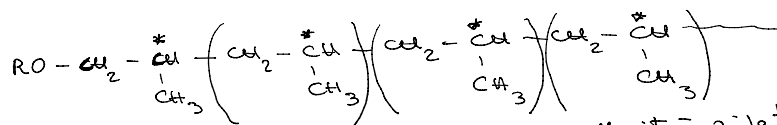
Propagazione
Addizione



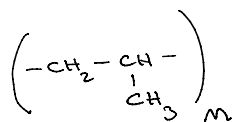
Transfer



L'inserimento è ordinato e di tipo TESTA-CODA



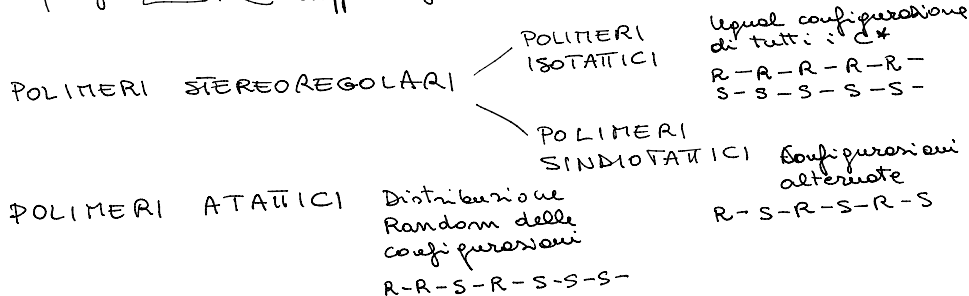
unità ripetitive
quello che "resta"
dell'alchene
dopo l'inserimento
nella catena



polipropilene
(PP)

nome del monomero

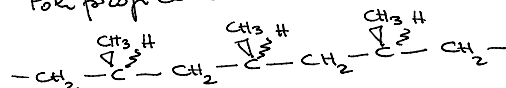
propil**ENE** - doppio legame



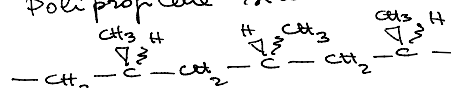
I polimeri stereoregolari hanno una temperatura di transizione vetrosa più alta di quelli atattici

Liquido
solido
Comune ↑ solido
Vetroso AT

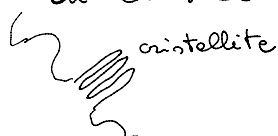
Poli**propilene** isotattico



Poli**propilene** sindiotattico



Polimeri stereoregolari hanno un certo grado di cristallinità (cristalliti)



I polimeri stereoregolari si ottengono utilizzando particolari catalizzatori (ZIEGLER-NATTA $\text{TiCl}_4 \text{ Al}(\text{C}_2\text{H}_5)_3$) meccanismo ionico di coordinazione.

MECCANISMO CATALONICO

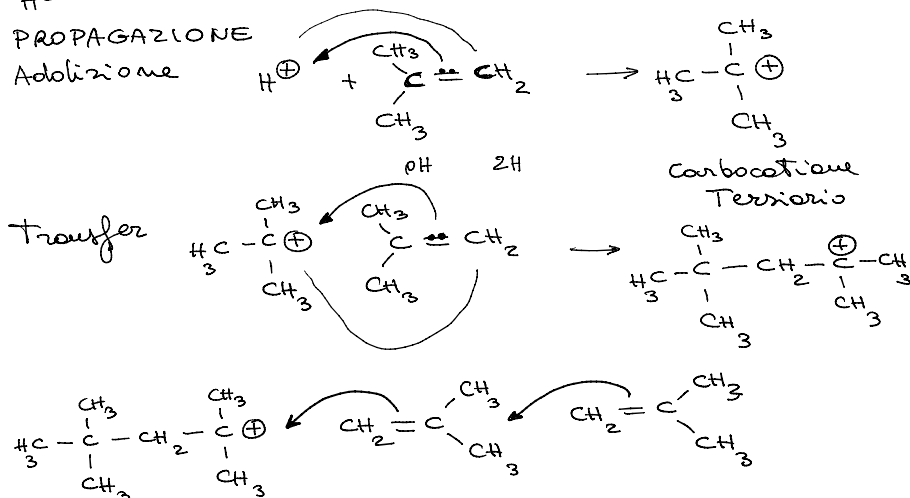
INIZIAZIONE

INIZIAZIONE
Come iniziatore si utilizza un acido protico (H_2SO_4)
oppure un acido di Lewis (BF_3 , $AlCl_3$)

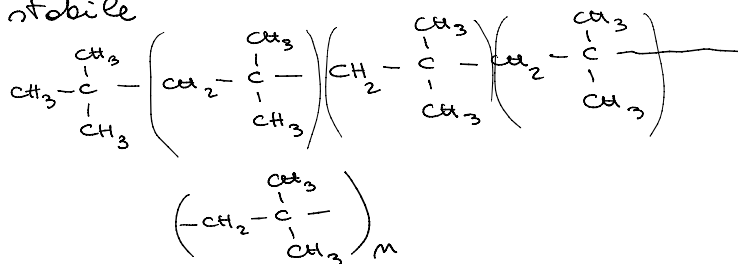


PROPAGAZIONE

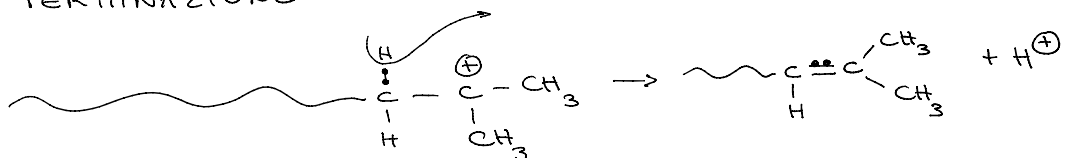
Adozione



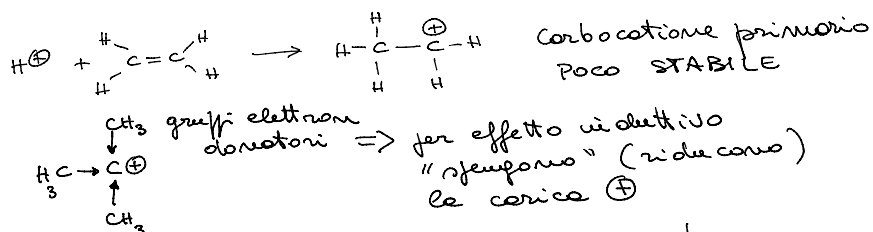
Inserimento ordinato TESTA-CODA che
permette di formare il carbocatione più
stabile



TERMINAZIONE



Per di' ho scelto $\begin{array}{c} \text{H}^3 \\ | \\ \text{C} \\ | \\ \text{H}^2 \end{array} = \text{CH}_2$ e non $\begin{array}{c} \text{H} \\ | \\ \text{C} \\ | \\ \text{H} \end{array} = \text{C} \begin{array}{c} \text{H} \\ | \\ \text{H} \end{array}$?



Per il meccanismo cationico vanno bene
alcheni con gruppi sostituenti: *electron-donatori*

MECCANISMO ANIONICO

INIZIAZIONE

INIZIAZIONE
Come iniziatore si utilizzano BASI FORTI
(N_2H_2^- o $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{Li}$)

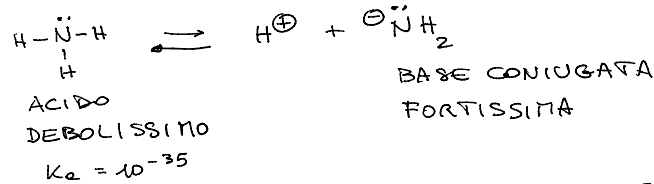


IONE AMMIDURO.
AMMONIURO

BASE FORTISSIMA
(più forte NaOH , KOH)



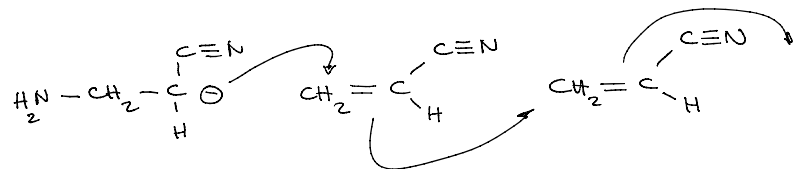
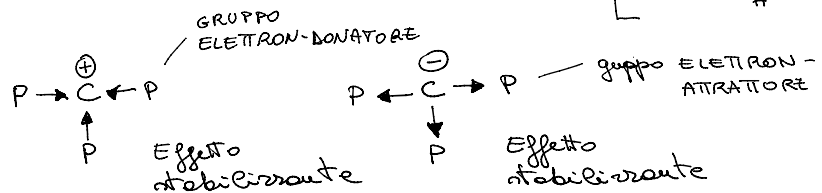
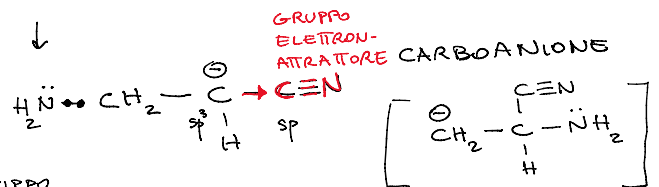
$\sim \text{NH}_2$ AMMONIURIO (più forte NaOH , KOH)



PROPAGAZIONE



ACRILONITRILE

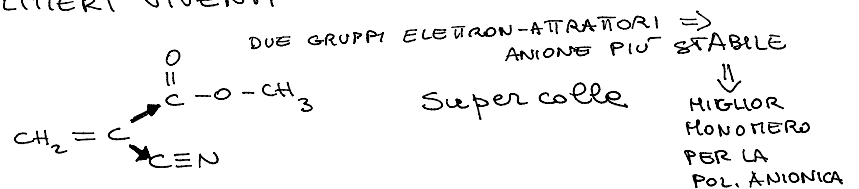


Inserimento TESTA-CODA (permette la formazione dell'anione più stabile)

TERMINAZIONE

Non esiste nella polimerizzazione anionica

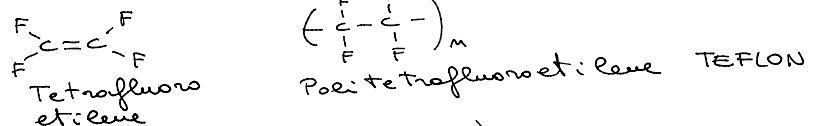
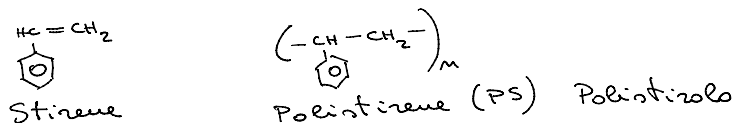
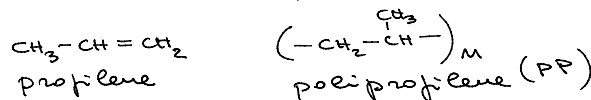
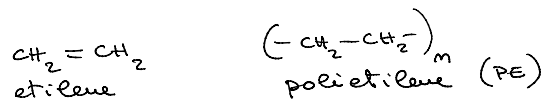
POLIMERI VIVENTI



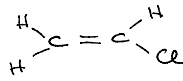
α -CIANO ACRILATO DI METILE

Monomero

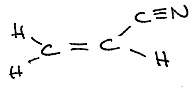
Polimero



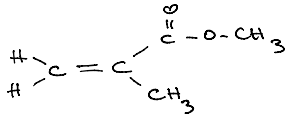
Tetrafluoro
etilene



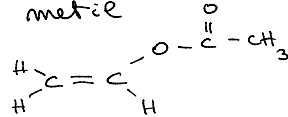
cloruro di
vinile



Acilonitrile

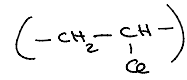


Metil Acrilato di
metile

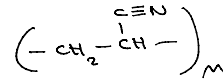


Acetato di
vinile

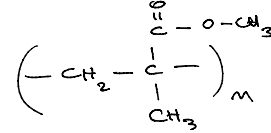
Poli tetrafluoro...



Poli vinil cloruro (PVC)

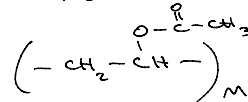


Poli acilonitrile ORLON



PLEXIGLAS

Poli Metil
Acrilato di Metile



Poli vinil acetato

Poli acetato di vinile